CHEMISTRY MCQ'S- 1 **mm.5**

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ACTIVITY 1

Aim: to classify the following reactions and observe them on the basis of 1) combination reaction
2) Decomposition reaction
3) Displacement reaction
4) double displacement reaction

(Q1) Action of water on quick lime is an example of:

1. Combination reaction
2. Displacement reaction
3. Decomposition reaction
4. Double displacement reaction

(Q2) action of heat on ferrous sulphate crystals is an example of:

1. Combination reaction
2. Displacement reaction
3. Decomposition reaction
4. Double displacement reaction

(Q3) Reaction of copper sulphate solution with iron fillings is an example of:

1. Combination reaction
2. Displacement reaction
3. Decomposition reaction
4. Double displacement reaction

(Q4) Reaction of barium chloride and sodium sulphate solution is an example of:

1. Combination reaction
2. Displacement reaction
3. Decomposition reaction
4. Double displacement reaction

(Q5) The colour of solid barium sulphate and solid sodium chloride is:

1. **White ,white**
2. **White, colorless**
3. **Colorless, white**
4. **Colorless white**

(Q6) When ferrous sulphate is heated, the gases evolved are:

1. **Sulphur dioxide**
2. **Sulphur trioxide**
3. **Both of them**
4. **None of these**

(Q7) When ferrous sulphate is heated the compound of the iron formed is:

1. **Ferric oxide**
2. **Ferrous oxide**
3. **Magnetite**
4. **All of these**

(Q8) When quick lime reacts with water the product so formed is:

1. **Quick lime**
2. **Slaked lime**
3. **Water**
4. **Lime water**

(Q9) Copper sulphate solution is also known as:

1. **Blue vitriol**
2. **Green vitriol**
3. **Milk of magnesia**
4. **Copper acid**

(Q10) The reaction of quick lime and water is:

1. **Endothermic**
2. **Exothermic**
3. **Both of them**
4. **None of these**